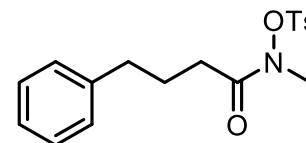
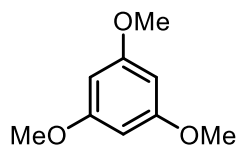


$$wt\% = \frac{m(std) \times M_w(5) \times \text{molar ratio} \times P(std)}{m(5) \times M_w(std)} = 99\%$$

m(**4**) = 34.4 mg; M_w(**4**) = 347.44 g/mol

m(std) = 16.5 mg; M_w(std) = 168.19 g/mol

P(std) = 98%; molar ratio = (3.07/3)/(3/3) = 1.02



4

